Introduction of Matter In Our Surroundings

Matter In Our Surroundings is the lesson totally about the particles present in everywhere and the solid, gas ,liquid and more.

Matter

Anything which occupies space or mass is also known as **matter**.

Particles

1. Particles are very small in size it also has mass.

Characteristics Matter In Our Surroundings

1. Particles also have space between them
2. these are continuously moving
3. it attract each other.

Solids Matter In Our Surroundings

1. Have definite shape also distinct boundaries and its volume is fixed.
2. Maintain their shape.
3. they are rigid and have negligible compressibility.

Liquid

1. it has fixed volume.
2. take up their shape of the container.
3. they are rigid and shape.

Gases

1. Particles move randomly due to high speed
2. However, they diffuse very fast in other gases
3. highly compressible

Change of states

1. Solid into vapor/vapor into solid - Sublimation
2. Solid into liquid - Melting
3. Liquid into Vapour - evaporation
4. Liquid into solid - freezing
5. Vapour into liquid - Condensation

Sublimation

1. Conversion of solid-state into a gaseous state and vice versa without passing through the liquid state
2. Occurs due to high vapour pressure of solid state.

Melting

1. conversion of solid to liquid state by the process of heat.
2. increase in temperature can cause increse in kinetic energy. this makes the particles move freely and one one point it start change as liquid.
3. the temperature at which the solid start to melt to became liquid is known as boiling point.

Evaporation

1. So,Convertion of liquid into vapour without reaching the boiling point.
2. supply of heat energy cause fast movement of particles. because of this at one point they are free from forces of attraction of each other and change to vapor.

Evaporation

Freezing

1. conversion of liquid state into solid-state at 0 Celsius degree.
2. Lowering the temperature decreases the kinetic energy of particles so, they start moving slowly and at one point change into a solid state.
3. The temperature at which liquid changes to solid at atmospheric pressure is its freezing point.

Condensation

1. conversion of the gaseous state into a liquid state.
2. occurs when the pressure of vapor becomes equal to the maximum pressure of the liquid at that temperature